# Question 2

- 2(a)(i) An acid is a substance which when dissolved in water produces H\* ions as the only positive ions.
- 2(a)(ii) A salt is the substance produces when some or all of the replaceable hydrogens of an acid have been replaced by a metal cation or the ammonium ion.
- 2(b)(i) The solution was alkaline and red litmus turns blue in alkaline conditions while blue litmus is unaffected by alkaline conditions.
- 2(b)(ii) Yes.
- 2(c)(i) Inference #1: Fe<sup>2+</sup>((aq) was formed. Inference #2: H, gas was produced.
- 2(c)(ii) It has a high melting point and is a good conductor of electricity.
- 2(c)(iii)  $Fe_{(s)} + H_2SO_{4(aq)} \rightarrow FeSO_{4(aq)} + H_{2(a)}$
- 2(c)(iv) Test: Bubble the gas through Ca(OH)<sub>2(aq)</sub> i.e. lime water. Observation: A white precipitate is initially formed.
- 2(c)(v)  $2NaHCO_{3(c)} + H_2SO_{4(aq)} \rightarrow Na_2SO_{4(aq)} + 2CO_{2(g)} + 2H_2O_{(i)}$ .

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