

Chemistry Modules 5 and 6

Practice exam marking guidelines

Section I—Solutions to multiple-choice questions

- 1 D. The activation energy is always positive. The value for the reverse reaction is $1370 + 572 = 1942 \text{ kJ mol}^{-1}$.
- 2 B. The free energy formula is $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$. A reaction is spontaneous if $\Delta G^\circ < 0$. For the value of ΔG° to be less than zero, the value of ΔH° has to be greater than $T\Delta S^\circ$ if both ΔH° and ΔS° are negative.
- 3 A. The graph shows two reactants and one product. One reaction drops at half the rate of the other, so the ratio of reactants is 2:1, which matches with option A.
- 4 C. Calcium chloride is an ionic solid. When mixed with water, the ionic bonds are broken as the ions dissociate. The hydrogen bonds between water molecules also break as bonds form between the ions and water molecules.
- 5 D. The superscript on the hydroxide ions is 3, confirming that the formula of aluminium hydroxide should be $\text{Al}(\text{OH})_3$. When it dissociates, the coefficient for the hydroxide ions should be 3.
- 6 A. The amount of NOCl drops by 0.6 mol. Therefore, the amount of NO should increase by 0.6 and the amount of Cl_2 will increase by half as much, 0.3 mol. The total number of moles will therefore be $1.0 + 0.6 + 0.3 = 1.9 \text{ mol}$.
- 7 B. The original equation has been reversed and halved. The new value of K_{eq} will be the square root of the reciprocal of the original value, $\frac{1}{\sqrt{0.065}} = 3.92$.
- 8 B. HCO_3^- is monoprotic as it can donate the hydrogen it contains to form CO_3^{2-} . It can also gain a proton to form H_2CO_3 ; hence, it is also amphiprotic.
- 9 C. The lowest pH corresponds to the strongest acid. Sulfuric acid is a strong acid and it is diprotic. Therefore, its pH will be lower than that of the monoprotic HCl or the weak acid, ethanoic acid.
- 10 D. $n\{\text{oxalic acid}\} = \frac{4.7}{126} = 0.037 \text{ mol}$
 $c = \frac{n}{V} = \frac{0.037}{0.25} = 0.15 \text{ mol L}^{-1}$

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Tobias Bleicker



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