Chem 102 Unit 1- Kinetics: Rates and Mechanisms of Chemical Reactions (Chapter 14).

Focusing On Reaction Rate (14.1)

Chemical kinetics is the study of the rates and mechanisms of reactions.

Reaction Rate: Refers to how fast or slow the concentration of a reactant or product changes per unit time. Different reactions have different rates.

Figure 14.1, p. 575)

Rate = change in concentration / time

=  $\Delta$ (concentration) /  $\Delta$ (time)

Can be: (i) increase in the concentration of product per unit time (ii) decrease in the concentration of reactant per unit time. Reaction mechanism refers to the reaction steps involved in going from reactants to products.

The reaction rate is proportional to the number of collisions, depends on the concentration of reactants.

Figures 14.2 & 14.3, p. 576; 14.4, p. 577.

The following factors affect the rate of reaction and can be controlled: concentration, physical state, temperature, frequency of collisions and addition of a catalyst.

Summary of Section 14.1, p. 577

Expressing the Reaction Rate (14.2)

Consider the reaction: A (reactant) ----→ B (product)

Rate = - change in concentration of A / change in time

Rate =  $-\Delta[A]/\Delta[t]$  or =  $+\Delta[B]/\Delta[t]$ 

Note: The plus (+) sign is usually understood but not shown

# **Chemistry 102 Lecture Notes**

South Dakota School of Mines and Technology

# **Chemistry 102 Lecture Notes:**

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