

CHEM 20024 Chemistry for Engineers  
Practice Problems, Chapters 7-10

**Direction/s:**

All questions are multiple choice in format. Copy the questionnaire and encircle the letter of your best answer. A periodic table and table of bond energies (if needed) are provided in our reference textbook.

- Rank the following elements below in order of increasing atomic radius (smallest to largest):  
Mg, Na, P, Si, Ar  
a) Mg, Na, P, Si, Ar      b) Ar, P, Si, Mg, Na      c) Si, P, Ar, Na, Mg      d) Na, Mg, Si, P, Ar
- Which one of the following atoms has the largest effective nuclear charge,  $Z_{eff}$ ?  
a) C      b) N      c) O      d) F
- Which of the following elements has the highest ionization energy?  
a) Na      b) Li      c) Cl      d) Ca
- When a metal and a nonmetal react, which one tends to lose electrons?  
a) the metal      b) the nonmetal      c) either      d) kebs
- Which one of the following ions has the largest radius?  
a)  $\text{Na}^+$       b)  $\text{Al}^{3+}$       c)  $\text{P}^{3-}$       d)  $\text{F}^-$
- Which will form an acidic solution when dissolved in water?  
a)  $\text{Li}_2\text{O}$       b)  $\text{Na}_2\text{O}$       c)  $\text{CO}_2$       d)  $\text{MgO}$
- The best Lewis dot structure of the NO molecule has  
a) an unpaired electron on the nitrogen atom  
b) an unpaired electron on the oxygen atom  
c) a triple bond between N and O  
d) an expanded octet on N
- The type of compound that is most likely to contain an ionic bond is  
a) one that is composed of a metal and a nonmetal  
b) a solid metal  
c) one that is composed of only nonmetals  
d) one that is nonpolar
- Which compound has the lowest lattice energy?  
a) NaCl      b) NaBr      c)  $\text{MgCl}_2$       d)  $\text{MgBr}_2$
- The lattice energy of an ionic solid is greater if  
a) the ions have large sizes  
b) the charges of the ions are high  
c) the distance between the ions is large  
d) the distance between the ions is large and the charges are low
- Which type of bond is the longest?  
a) single      b) double      c) triple      d) all are the same length

# Chemical Eng Practice Volume 12 Indexes

**Rachel Sandford**



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