

Exercise 7-1: Calculations for Solutions

Use the formula for concentration of a solution to solve the following problems for the desired quantity. For this assignment you will need to print this page and complete each problem using a pencil and your calculator. Concentration of solution = $\frac{\text{amount of solute}}{\text{amount of solution}}$

a. A saturated solution will form when 357 grams of sodium chloride is dissolved in one liter of water at 0 °C. How many grams of sodium chloride are necessary to form 2.5 liters of saturated solution?	Solve using formula for amount of solute $\text{amount of solute} = \text{Concentration} \times \text{amount of solution}$ $= 357 \frac{\text{grams}}{1 \text{ L}} \times 2.5 \text{ L} = 892.5 \text{ g} = \boxed{892.5 \text{ g}}$
b. A saturated solution will form when 1220 grams of silver nitrate is dissolved in one liter of water at 0 °C. How many grams of silver nitrate are necessary to form 0.500 liters of saturated solution?	Solve for amount of solute $\text{amount of solute} = \text{Concentration} \times \text{amount of solution}$ $= 1220 \frac{\text{g}}{\text{L}} \times 0.500 \text{ L} = \boxed{610 \text{ g}}$
c. A student is told to make a solution of sodium chloride that has a concentration of 15 grams per liter. How many liters of solution can they make if they have 45 grams of sodium chloride?	Solve for amount of solution $\text{amount of solution} = \frac{\text{amount of solute}}{\text{Concentration of solution}}$ $= \frac{45 \text{ g}}{15 \frac{\text{g}}{\text{L}}} = \boxed{3.0 \text{ L}}$
d. How many moles of sugar is required to make 1.5 liters of a 2.0 molar solution? How many sugar molecules is this?	Molarity = $\frac{\text{Number moles of solute}}{\text{Liters of solution}}$ $\text{moles of solute} = \text{Molarity} \times \text{Liters of solution}$ $= 2.0 \frac{\text{moles}}{\text{L}} \times 1.5 \text{ L} = \boxed{3.0 \text{ moles}}$ $\text{Number of molecules} = 3 \text{ moles} \times 6.02 \times 10^{23} \frac{\text{molecules}}{\text{mole}}$
e. Calculate the Molarity of a solution prepared by dissolving 0.20 mol of sodium chloride in enough water to make 1.6 liters of solution.	Molarity = $\frac{\text{Number moles of solute}}{\text{Liters of solution}}$ $= \frac{0.20 \text{ mol}}{1.6 \text{ L}} = \boxed{0.125 \text{ M}}$
f. Calculate the Molarity of a solution when 1.5 mol of a substance is dissolved in water to form 0.50 liters of solution.	Molarity = $\frac{1.5 \text{ mol}}{0.50 \text{ L}} = \boxed{3.0 \text{ M}}$
g. What volume (in liters) of solution is required to provide 2 moles of substance from a solution with a Molarity of 0.50 mol per liter of solution?	Solve for Liters of solution $\text{Liters of solution} = \frac{\text{Number moles of solute}}{\text{Molarity}}$ $= \frac{2 \text{ moles}}{0.50 \frac{\text{moles}}{\text{L}}} = \boxed{4.0 \text{ L}}$

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