

# Calculations in Analytical Chemistry

## The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is,  $N_A = 6.022 \times 10^{23}$ .

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

Element	Molar mass per element	Number of atoms	Mass
C	12 g/mol	6	72 g/mol
H	1 g/mol	12	12 g/mol
O	16 g/mol	6	96 g/mol
Molar Mass			180 g/mol

## Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

1. Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 g_{HBz} \times \frac{1 mol_{HBz}}{122.1 g_{HBz}} = 0.0164 mol_{HBz}$$

To obtain the number of millimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 g_{HBz} \times \frac{1 mmol_{HBz}}{0.1221 g_{HBz}} = 16.4 mmol_{HBz}$$

2. What is the mass in grams of  $Na^+$  (22.99 g/mol) in 25.0 g of  $Na_2SO_4$  (142.0 g/mol)?

$$mass_{Na^+} = 25.0 g_{Na_2SO_4} \times \frac{1 mol_{Na_2SO_4}}{142.0 g_{Na_2SO_4}} \times \frac{2 mol_{Na^+}}{1 mol_{Na_2SO_4}} \times \frac{22.99 g_{Na^+}}{1 mol_{Na^+}}$$

$$mass_{Na^+} = 8.10 g_{Na^+}$$

# Calculations In Analytical Chemistry

**Leicester Forsyth Hamilton, Stephen  
Gershom Simpson, David W. Ellis**



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