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Ying-Ying Zheng



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web oct 15 2014 1 answer sorted by 6 here are the three equations describing the reduction of manganese and concurrent oxidation of whatever substrate may be present under basic neutral and acidic conditions respectively $\text{Mn}^{7+} \rightarrow \text{Mn}^{4+}$ $\text{Mn}^{7+} \rightarrow \text{Mn}^{2+}$ $\text{Mn}^{7+} \rightarrow \text{Mn}^{2+}$ basic $\text{Mn}^{7+} \rightarrow \text{Mn}^{4+}$ $\text{Mn}^{7+} \rightarrow \text{Mn}^{2+}$ $\text{Mn}^{7+} \rightarrow \text{Mn}^{2+}$ basic

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web mar 15 2019 what is the mechanism of toluene oxidation by KMnO_4 it is quite simple reaction that KMnO_4 can oxidize

benzylic hydrogen in toluene and synthesize benzoic acid however i cannot find exact mechanism of reaction in

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web description treatment of an alkylbenzene with potassium permanganate results in oxidation to give the benzoic acid notes the position directly adjacent to an aromatic group is called the benzylic position the reaction only works if there is a hydrogen attached to the carbon examples

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web may 1 2004 the oxidation of toluene by aq KMnO_4 under hydrodynamic cavitation was taken as a model reaction and various parameters have been optimized the oxidation of toluene by aq KMnO_4 gives benzoic acid scheme 1 KMnO_4 in turn is reduced to MnO_2 the reaction did not give any other byproduct