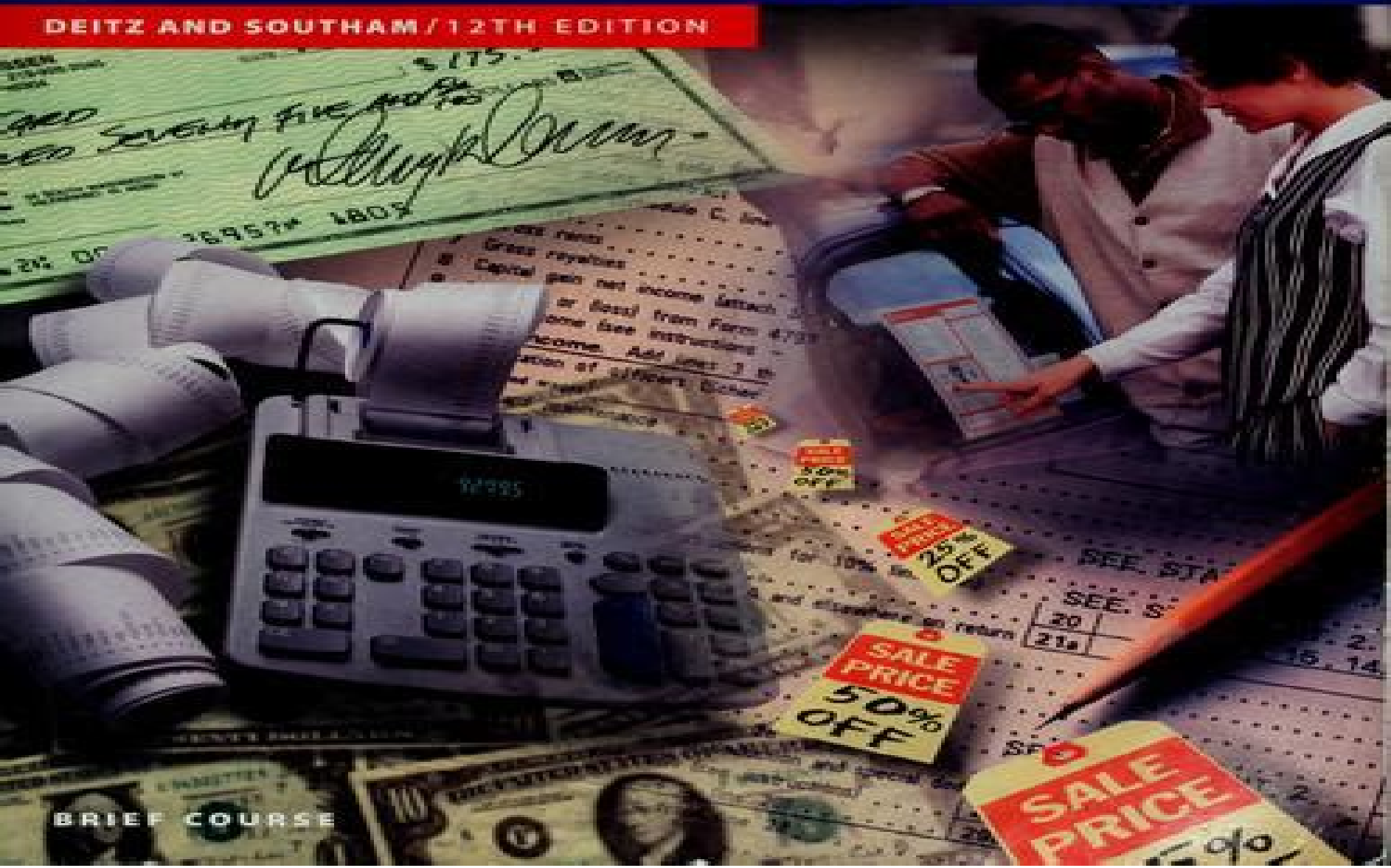


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web solution average atomic mass of li 6 941 amu x 6 0151 amu 1 x 7 0160 amu 6 941 1 0009x 7 0160 1 0009x 0 075 x 0 075 x 0 075 corresponds to a natural abundance of 6li of 7 5 percent the natural abundance of 7li is x 0 925 or 92 5 percent 3 7 the unit factor required is 6 022

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web a molar mass li₂co₃ 2 6 941 g 12 01 g 3 16 00 g 73 89 g b molar mass cs₂ 12 01 g 2 32 07 g 76 15 g c molar mass chcl₃ 12 01 g 1 008 g 3 35 45 g 119 37 g d molar mass c₆h₈o₆ 6 12 01 g 8 1 008 g 6 16 00 g 176 12 g e molar mass kno₃ 39 10 g 14 01 g 3 16 00 g 101 11 g

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web it is a small step from mole mass calculations to mass mass calculations if we start with a known mass of one substance

in a chemical reaction instead of a known number of moles we can calculate the corresponding masses of other substances in the reaction

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web jul 4 2022 to calculate the mass of titanium metal that you can obtain multiply the number of moles of titanium by the molar mass of titanium 47.867 g/mol moles ti mass ti times molar mass ti $4.12 \text{ mol ti} \times \frac{47.867 \text{ g ti}}{1 \text{ mol ti}} = 197 \text{ g ti}$ notag

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web 41 42 chemistry ch 3 mass relationships in chemical reactions molecular mass of NO_2 atomic mass of n 2 atomic mass of o 14 01 amu $2 \times 16.00 \text{ amu} = 32.00 \text{ amu}$ atomic mass of N_2O_5 2 atomic mass of n 5 atomic mass of o 2 14 01 amu $5 \times 16.00 \text{ amu} = 80.00 \text{ amu}$ example 3 2 exercise 3 3