

Exploring the Best

# AI



# BUSINESS IDEAS

A Complete Guide

# Beginner Ai Business Ideas Guide With Low Investment

**VM Jensen**



## **Beginner Ai Business Ideas Guide With Low Investment:**

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point to keep in mind is that  $n$  can be moles of any gas since one mole always contains  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as  $6.02 \times 10^{23}$  particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always  $6.02 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the SI unit of cubic metres per mole  $\text{m}^3 \text{mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure STP is  $22.4 \text{ L}$

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web at standard temperature and pressure STP the molar volume  $V_m$  is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass  $M$  by mass density  $\rho$  molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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web after defining the gas phase properties of gases such as temperature pressure and volume are discussed the relationships between these properties are explained at a molecular level and simple models are introduced that allow the

various gas laws to be derived from first principles finally the collisional behavior of gases is used to

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web this volume is known as the molar gas volume  $V_m$  the volume of a gas may be calculated from its number of moles using volume of gas moles  $V_m$  rearranging this moles of gas  $\frac{V}{V_m}$

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web jun 8 2023  $V = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{ K} = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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