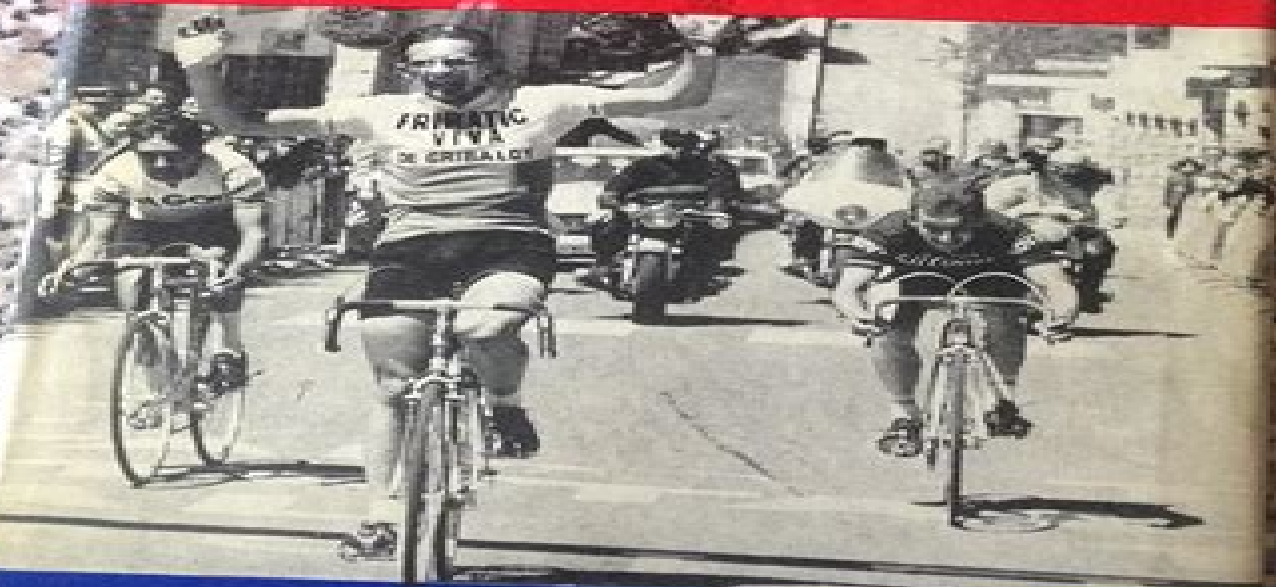


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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+\text{aq} + \text{oh}^-\text{aq} \rightarrow \text{h}_2\text{o}(\text{l})$ the spectator ions in $\text{koh}(\text{aq})$ and $\text{hno}_3(\text{aq})$ are k^+aq and no_3^-aq how to identify and list the

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web the ag^+aq and cl^-aq ions become $\text{agcl}(\text{s})$ but the na^+aq ions and the no_3^-aq ions stay as na^+aq ions and no_3^-aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction

the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be
 $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq}$
 $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq}$ these spectator ions are not actively involved in the reaction the existence of
spectator ions affects the debye length