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web video answer so now we ll talk about problems 58 from chapter four this problem whereas how do isotopes differ and how are they similar isotopes for the same element so isotopes differ in the number of neutro

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web the isotopes of a given element always contain the same number of protons and therefore occupy the same place on the periodic table however because isotopes contain different numbers of neutrons each isotope has a unique atomic mass

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almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web may 16 2018 explanation an atom is characterized as a particular element by the number of positively charged massive nuclear particles i e z $z = 1$ we got hydrogen $z = 2$ we got helium $z = 3$ lithium $z = 47$ we got silver the nucleus can also contain neutrons massive nuclear particles of zero charge

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web atoms of the same element with different numbers of neutrons are called isotopes close isotope atoms of an element with the same number of protons and electrons but different numbers of neutrons

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same

atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope s mass number

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