

UGC RATES: For New Creators

Single Concept

A SINGLE UGC CONCEPT FOR USE ON ORGANIC OR PAID SOCIAL MEDIA.

\$150 - 200

Additional Hooks/CTAs

ADDITIONAL HOOKS OR CTA'S (START + END) THAT ENABLES BRANDS TO MAKE VARIATIONS.

\$50 EACH

Usage Rights

IF THE BRAND WANTS TO EXTEND THE USAGE RIGHTS BEYOND WHAT YOU INCLUDE IN YOUR BASE RATE.

30-50% BASE RATE

Raw Footage

UNEDITED FOOTAGE (ALL OF YOUR UMS, AHS AND SWEAR WORDS).

30-50% BASE RATE

Whitelisting

ALLOWING THE BRAND TO BOOST A VIDEO FROM YOUR ACCOUNT, WITH BRAND SPEND

30% BASE RATE PER MONTH

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{Au} + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal

convert to basic solution $4\text{CN}^- + 2\text{Au} + \text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{OH}^-$ 3 the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}_2$ NO_2 choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and NO_2 c h and

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