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to generate creative outputs New applications for creative tasks are rapidly evolving and new tools are constantly being developed with much greater optimal capabilities Importantly the success of implementing such tools for creative tasks is still heavily dependent on human supervision and input Therefore it is vital to understand and critically reflect on the nature of co creative processes between humans and AI This book addresses such issues and provides insights into how humans can augment their capabilities for generating creative and innovative outputs by successfully co creating with AI The book is intentionally divided into three main parts to allow for a comprehensive and holistic perspective on human and AI co creation for creative tasks The sections are divided as follows Part 1 Principles of AI and Creativity Part 2 Critical Issues on Artificial Co Creation and Part 3 Industry Specific Discussions Consequently the book provides a holistic insight on the topic covering various issues and perspectives and enabling an accessible read to a broad audience For example chapters cover examples across different industry sectors including music arts science and management Furthermore the book covers critical questions involving copyrights ethical concerns relationship with algorithms and context based issues Only by critically reflecting on the intrinsic issues of AI and learning how to work with it effectively for creative purposes will we be able to benefit from its full potential to augment human creative abilities in an appropriate manner This novel edited collection is an essential read for scholars working on the intersection of AI creativity arts and management

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b  $\text{Ba}^{2+}$  c  $\text{Br}_2$  d ca e  $\text{Ga}^{3+}$  f h 2 g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c  $\text{KClO}_3$  d  $\text{CH}_3\text{OH}$  2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe} + \text{Co} \rightarrow \text{Fe}^{2+} + \text{Co}^{2+}$  oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ag} + \text{Ni}^{2+}$  oxidation reduction 25  $\text{Cu} + \text{Pb} \rightarrow \text{Cu} + \text{Pb}^{2+}$  oxidation reduction 26  $\text{O}_2$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + \text{O}_5 \rightarrow 5\text{CO}_2 + \text{I}_2$  5  $\text{CO}_2 + \text{I}_2 \rightarrow 5\text{CO} + \text{I}_2$  5  $\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also 2  $\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr} \rightarrow \text{Sr}^{2+}$  oxidized reducing agent  $\text{O}_2 \rightarrow \text{O}^{2-}$  reduced ox ag 2  $\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li} \rightarrow \text{Li}^{+}$  oxidized red ag  $\text{S} \rightarrow \text{S}^{2-}$  reduced ox ag 3 2  $\text{Cs} + \text{Br}_2$

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