

The Development of Medical Techniques and Treatments

From Leeches to Heart Surgery

Martin Duke, M.D.



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RJ Alexander



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web a at atmospheric pressure helium can exist in all three phases as well as a supercritical fluid phase near absolute zero

choice b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure choice c

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web stant is the foundation of thermochemistry it allows us to describe heat flow in terms of the change in a state function enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if the volume of the system remains constant and thus no thermodynamic work is performed 5 b

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web thermochemical equations equations that represent both mass and enthalpy changes $H_2O(s) \rightarrow H_2O(l) \Delta H = 6.01 \text{ kJ/mol}$ this is

an endothermic process it requires 6.01 kJ to melt one mole of ice. ΔH_{fus} the enthalpy value will change if the number of moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy ΔU of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings $\Delta U = q + w$

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