

$$\left(P + \frac{a}{v^2}\right)(v - b) = RT$$

$$P = \left[ \frac{RT}{v - b} - \frac{a}{v^2} \right]$$

For reversible process

$$P_{\text{ext.}} = P_{\text{gas}}$$

$$W = - \int \left( \frac{RT}{v - b} - \frac{a}{v^2} \right) dv$$

So, process is not applicable only for irreversible process.

Hence Ans. A,B,C

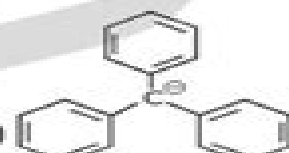
8. With respect to the compounds I-V, choose the correct statement(s).



- (A) The acidity of compound I is due to delocalization in the conjugate base.  
 (B) The conjugate base of compound IV is aromatic.  
 (C) Compound II becomes more acidic, when it has a -NO<sub>2</sub> substituent.  
 (D) The acidity of compounds follows the order I > IV > V > II > III.

Answer: A, B, C

Solution:

- (A)  is a conjugate base of compound I Which is stabilized by delocalisation or resonance.

- (B)  is a conjugate base of , which is an aromatic compound.

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