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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1 0 10 0 0585 correct answer b 1 explanation molality begin array l m left frac 1000 m 1000d mm right end array

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital m signifies solutions labelled with molar concentration a 1.0 M solution contains 1 mole of solute per litre of solution molality is the number of solvent moles per kilogram

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web 1 molarity mv mass molar mass $\times 0.4500 \text{ L} \times 0.825 \text{ g} \times 141.9579 \text{ g/mol} \times 0.0129 \text{ M} \times 2 \text{ molality } 0.825 \text{ g} \times 141.9579 \text{ g/mol} \times 0.00581158 \text{ mol} \times 0.00581158 \text{ mol} \times 0.4500 \text{ kg} \times 0.0129 \text{ M} \times 3 \text{ mole fraction } n_2 \text{ hpo}_4 \times 0.825 \text{ g} \times 141.9579 \text{ g/mol} \times 0.00581158 \text{ mol} \times 2 \times 0.4500 \text{ g} \times 18.015 \text{ g/mol} \times 24.97918401 \text{ mol}$

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web apr 7 2014 molarity moles of solute liters of solution 8.422 a first convert 250 ml to liters 250/1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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